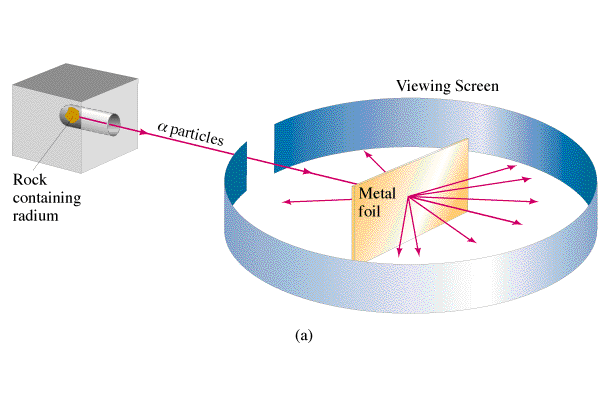
Atomic Models - Thomson and Rutherford

**Rutherford** – Discovered the nucleus by scattering alpha particles (2 protons, 2 neutrons bound together) off of gold foil.



|  |  |
| --- | --- |
| Relationship between energy voltage and charge:    Ve = Voltage (V)  q = Charge (C)  Ep = Electrical Potential energy (J) | Voltage due to a point charge:    Ve = Potential near a point charge (V)  k = 8.99x109 Nm2/C2  q = Charge (C)  r = distance to charge (m) |
| Kinetic Energy:    Ek = Kinetic Energy (J)  m = mass (kg)  v = velocity (m/s) | Rutherford Nuclear Radius:    R - Nuclear radius (m)  Ro - Fermi Radius (1.20x10-15 m)  A - Mass # (#p +#n) |

Example 1: What is the radius of a Uranium 235 nucleus? (A = 235)

Example 2: What is the closest approach of an alpha particle going 2.6 x 106 m/s if it approaches a carbon nucleus head on?

Example 3: Through what potential must you accelerate an alpha particle to penetrate a Uranium (Z = 92) nucleus? (r = 7.4 fm) (1 fm = 1x10-15 m)